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CHAPTER 4

IMPERFECTIONS IN SOLIDS

PROBLEM SOLUTIONS

Vacancies and Self-Interstitials

4.1 The equilibrium fraction of lattice sites that are vacant in silver (Ag) at 700°C is 2×10^{-6} . Calculate the number of vacancies (per meter cubed) at 700°C. Assume a density of 10.35 g/cm³ for Ag.

Solution

This problem is solved using two steps: (1) calculate the total number of lattice sites in silver, N_{Ag} , using Equation 4.2; and (2) multiply this number by fraction of lattice that are vacant, 2×10^{-6} . The parameter N_{Ag} is related to the density, (ρ), Avogadro's number (N_A), and the atomic weight (A_{Ag} =107.87 g/mol, from inside the front cover) according to Equation 4.2 as

$$N_{Ag} = \frac{N_A \cdot \rho}{A_{Ag}}$$
$$= \frac{(6.022 \cdot 10^{23} \text{ atoms/mol})(10.35 \text{ g/cm}^3)(10^6 \text{ cm}^3/\text{m}^3)}{107.87 \text{ g/mol}}$$
$$= 5.78 \times 10^{28} \text{ atoms/m}^3$$

The number of vacancies per meter cubed in silver at 700°C, N_v , is determined as follows:

$$N_v = (2 \cdot 10^{-6})N_{Ag}$$

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